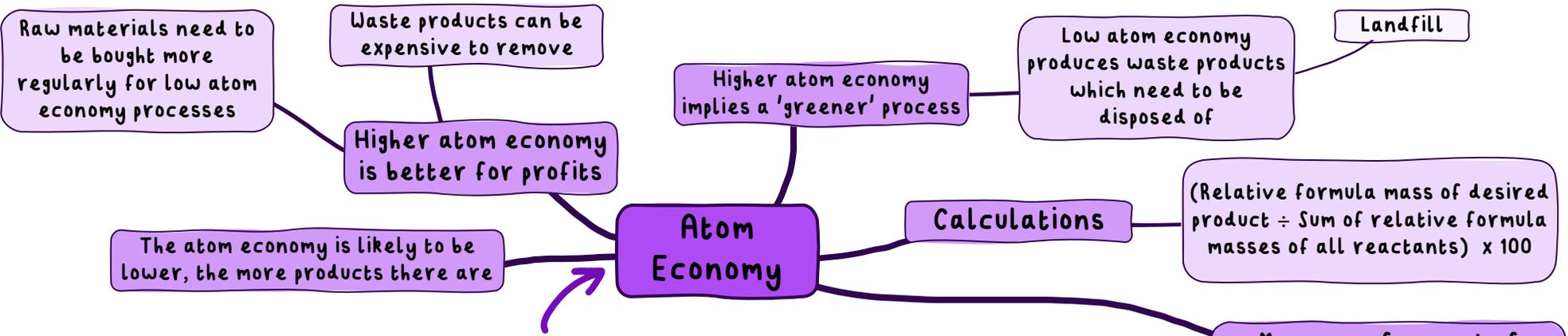


# 3.3 YIELD AND ATOM ECONOMY OF CHEMICAL REACTIONS (chemistry only)

## Atom Economy



Reversible reaction

E.g. Haber process

Incomplete reaction

## Percentage yield

Not always possible to obtain theoretical amount of product

Unexpected side reactions

Product lost when separated from reaction mixture

Calculations

$$\left( \frac{\text{Mass of actual yield}}{\text{Mass of theoretical yield}} \right) \times 100$$

**AQA**

